Chemistry
Moles practice

Calculate the molar mass of:

1. $NaHCO_3$ 2. $Mg(NO_3)_2$ 3. $(NH_4)_3PO_4$

4. Calculate the mass of:

   a) 0.948 mol of $NaHCO_3$

   b) 3.67 mol of $Mg(NO_3)_2$
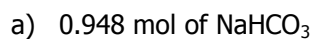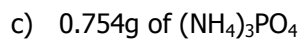
   c) 2.94 mol of $(NH_4)_3PO_4$

5. Calculate the number of moles in:

   a) 9.67g of $NaHCO_3$

   b) 38.4g of $Mg(NO_3)_2$

   c) 0.754g of $(NH_4)_3PO_4$
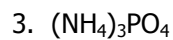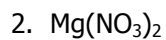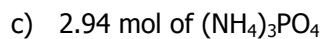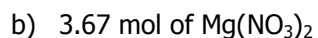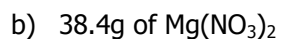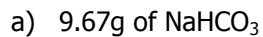
6. Calculate the percent composition of:

a) $NaHCO_3$ b) $Mg(NO_3)_2$ c) $(NH_4)_3PO_4$

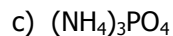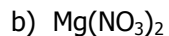7. Calculate the number of grams of oxygen in 25.0g of

   a) $NaHCO_3$

   b) $Mg(NO_3)_2$

   c) $(NH_4)_3PO_4$
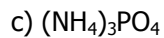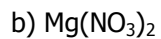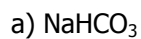
Moles - Problem Set I

1. Determine the molar mass of each of the following compounds:

    (a) $HClO_3$ _____   (b) $(NH_4)_2S$ _____

    (c) $Ca(NO_3)_2$ _____   (d) UF_6 _____

2. Calculate the number of grams contained in 0. 3 mol of each compound listed in question **one**.

    (a) _____ g   (b) _____ g

    (c) _____ g   (d) _____ g

3. Calculate the number of moles contained in 100.0 g of each compound listed in question **one**.

    (a) _____ mol   (b) _____ mol

    (c) _____ mol   (d) _____ mol

4. In 3.01×10^{24} molecules of water:

    (a) How many moles of water are present ? _____ mol H_2O

    (b) How many grams of water are present ? _____ g H_2O